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11 chap 6 | Thermodynamics 07 || Heat of Reaction |  
Enthalpy Of Formation | Enthalpy Of Combustion | 90  
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As liquids absorb heat at their boiling points, the  
temperature remains constant while they vaporize.

true \_\_\_\_ Chapter 17 Thermochemistry 187

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188 Name \_\_\_\_ Date \_\_\_\_ Class \_\_\_\_ CHAPTER 17,

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Thermochemistry (continued) Use the heating curve for water shown below to answer Questions 5, 6, and 7.

## **section 17.1 the flow of energy heat and work**

Chapter 17 Thermochemistry Packet Answers 2. 0.100 g of H<sub>2</sub> and an excess of O<sub>2</sub> are compressed into a bomb calorimeter containing 1200 g of water. The temperature before the reaction is 25.00 °C, and after the reaction it goes to 27.16 °C.

## **Chapter 17 Thermochemistry Packet**

CHAPTER17, Thermochemistry (continued) Use the heating curve for water shown below to answer

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Questions 5,6, and 7. Heating Curve for Water Boiling point ~::>,------,1 ~ Melting E point ~ <--\_-----'l Heat supplied 5. Label the melting point and boiling point temperatures on the graph. 6. What happens to the temperature during melting and vaporization?

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Chapter 17 Thermochemistry 187 10. Complete the enthalpy diagram for the combustion of natural gas. Use the thermochemical equation in the first paragraph on page 517 as a guide. SECTION 17.3 HEAT IN CHANGES OF STATE (pages 520–526) This section explains heat transfers that occur during melting, freezing, boiling, and condensing.

## **SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505–510)**

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1. How much energy must be absorbed by 20.0 g of water to increase its temperature from 283.0 °C to 303.0 °C? 2. When 15.0 g of steam drops in temperature from 275.0 °C to 250.0 °C, how much heat energy is released?

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