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Nuclear fission and nuclear fusion - what exactly happens in these processes? GCSE Physics - Radioactive Decay and Half Life #35 NUCLEUS - Q VALUE OF NUCLEAR REACTION

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The phenomenon of radioactivity was discovered by Henry Becquerel in 1896. It may be defined as the phenomenon of spontaneous emission of high energy

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radiations (alpha rays, beta rays and gamma rays) from certain heavy elements by disintegration of their heavy nuclei. And the substances which emit such radiations are called radioactive substances.

~~RADIOACTIVITY & NUCLEAR REACTIONS~~
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Nuclear reactions are the reactions that include the change in the composition of the nuclei of the reactants elements and the formation of new nuclei when the nuclei of the reacting atoms interact, The nuclear reactions can be classified into four types :
Natural transformation of elements (Natural Radioactivity)

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~~Radioactivity, Nuclear reactions (Natural transformation ...~~

Ionizing radiation is the most harmful because it can ionize molecules or break chemical bonds, which damages the molecule and causes malfunctions in cell processes. Types of radiation differ in their ability to penetrate material and damage tissue, with alpha particles the least penetrating but potentially most damaging and gamma rays are most penetrating.

17.11: Radioactivity in Medicine The field of nuclear medicine has expanded greatly in the last twenty years.

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~~17: Radioactivity and Nuclear Chemistry — Chemistry LibreTexts~~

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Radioactivity is the spontaneous emission of particles or radiation or both at the same time. These particles

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and radiation come from the decay of certain nuclides that form them. They disintegrate due to a fix in their internal structure. Radioactive decay occurs in unstable atomic nuclei. That is, those that do not have enough binding energy to hold the nucleus together.

~~What Is Radioactivity? Definition and Types – Nuclear Energy~~

Radioactivity And Nuclear Reaction - Displaying top 8 worksheets found for this concept. Some of the worksheets for this concept are Nuclear reaction work answer key, Nuclear reactions review work, Chm152II nuclear chemistry summer work, Radioactivity and

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~~Radioactivity And Nuclear Reaction Worksheets
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Since first-order reactions have already been covered in detail in the kinetics chapter, we will now apply those concepts to nuclear decay reactions. Each radioactive nuclide has a characteristic, constant half-life ($t_{1/2}$), the time required for half of the atoms in a sample to decay.

~~3.1: Nuclear Chemistry and Radioactive Decay~~

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~~Chemistry ...~~

Chapter 13 The Atomic Nucleus And Radioactivity -
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Radioactive decay (also known as nuclear decay, radioactivity, radioactive disintegration or nuclear

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disintegration) is the process by which an unstable atomic nucleus loses energy by radiation. A material containing unstable nuclei is considered radioactive .

~~Radioactive decay - Wikipedia~~

Unstable atomic nuclei will spontaneously decompose to form nuclei with higher stability. The decomposition process is called radioactivity. The energy and particles which are released during the decomposition process are called radiation. When unstable nuclei decompose in nature, the process is referred to as natural radioactivity.

~~Quick Review of Radioactivity and Radiation~~

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Nuclear Chemistry: Half-Lives and Radioactive Dating;
Nuclear Chemistry: Half-Lives and Radioactive Dating.
Scientists look at half-life decay rates of radioactive isotopes to estimate when a particular atom might decay. A useful application of half-lives is radioactive dating. This has to do with figuring out the age of ancient things.

~~Nuclear Chemistry: Half-Lives and Radioactive Dating~~
PSc.2.3 Understand the role of the nucleus in radiation and radioactivity. PSc.2.3.1 Compare nuclear reactions including: alpha decay, beta decay, and gamma decay; nuclear fusion and nuclear fission.
□ Compare the characteristics of alpha and beta

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particles and gamma rays – composition, mass, penetrability.

~~Radioactivity – MS. SMITH'S CLASS~~

Half-life, in radioactivity, the interval of time required for one-half of the atomic nuclei of a radioactive sample to decay (change spontaneously into other nuclear species by emitting particles and energy), or, equivalently, the time interval required for the number of disintegrations per second of a radioactive material to decrease by one-half.

~~half life | Definition & Facts | Britannica~~

Radioactivity is a spontaneous process in which atoms

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with unstable ratios of protons and neutrons in their nuclei decompose into forms with higher stability. Radioactivity is unique because atoms can change their identities by altering the number of protons in their nuclei. radioactivity alpha beta gamma

~~Radioactivity Concept Chemistry Video by Brightstorm~~

In general, radioactivity (known also as nuclear decay or radioactive decay) is a random process at the level of single atoms. Radioactivity occurs when an unstable atom loses energy by emitting ionizing radiation.

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~~Radioactivity — Nuclear Decay~~

Nuclear transmutation is the conversion of one chemical element or an isotope into another chemical element. Because any element (or isotope of one) is defined by its number of protons (and neutrons) in its atoms, i.e. in the atomic nucleus, nuclear transmutation occurs in any process where the number of protons or neutrons in the nucleus is changed.. A transmutation can be achieved either by

...

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